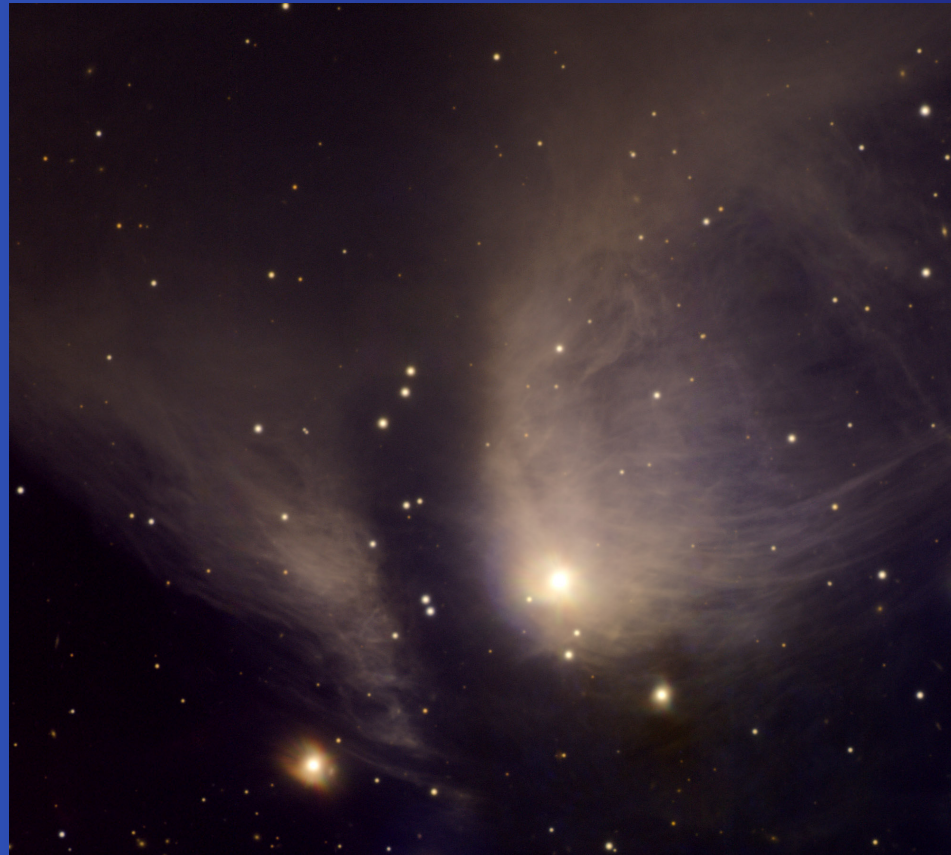


# Monte Carlo simulations on the formation of interstellar ice

Herma M. Cuppen, and Eric Herbst

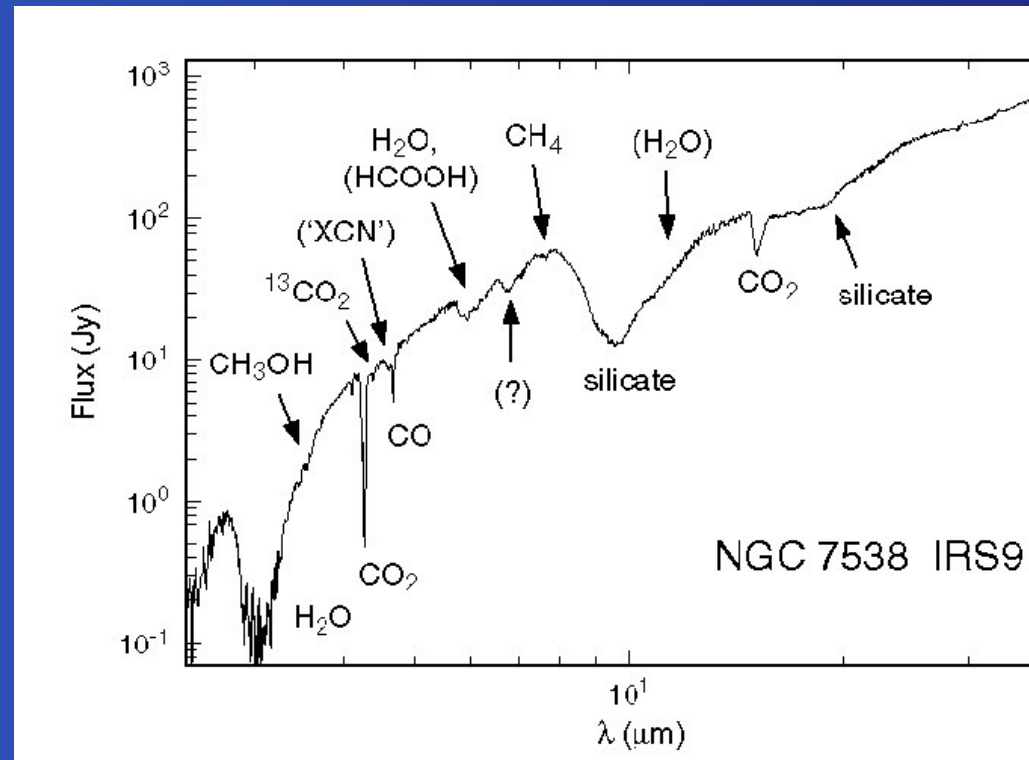
The Ohio State University, Columbus

# Molecular Cloud



RY Tauri

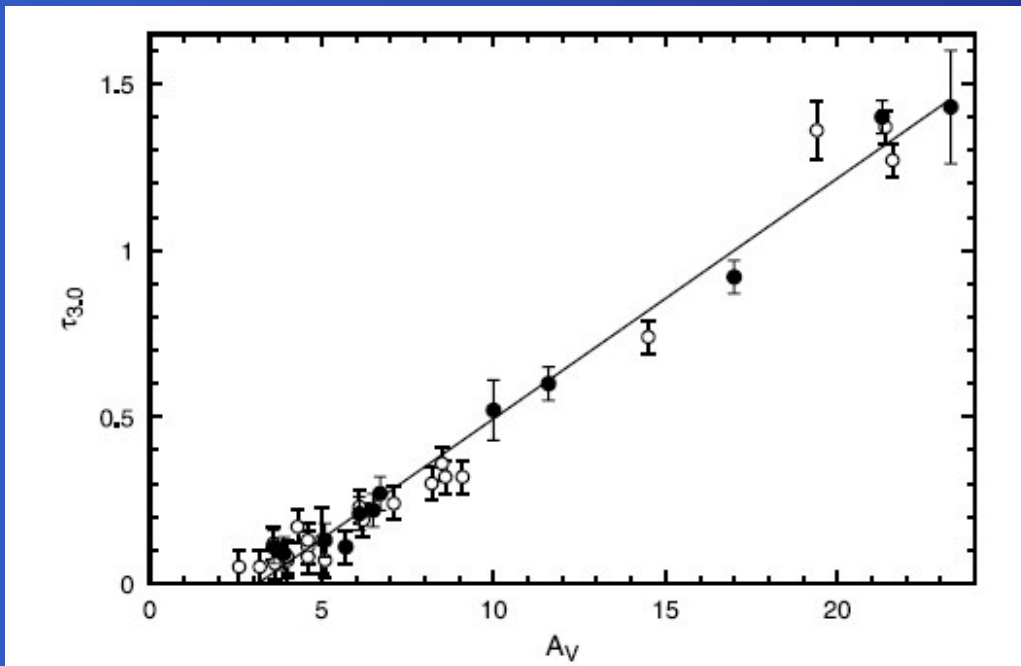
# IR spectrum towards a protostar



Dominant mantle species: water,  $\text{CO}_2$ ,  $\text{CO}$ ,  $\text{CH}_3\text{OH}$

# Interstellar ice vs extinction

$$I_V(x) = I_V(0)10^{-A_V/2.5}$$

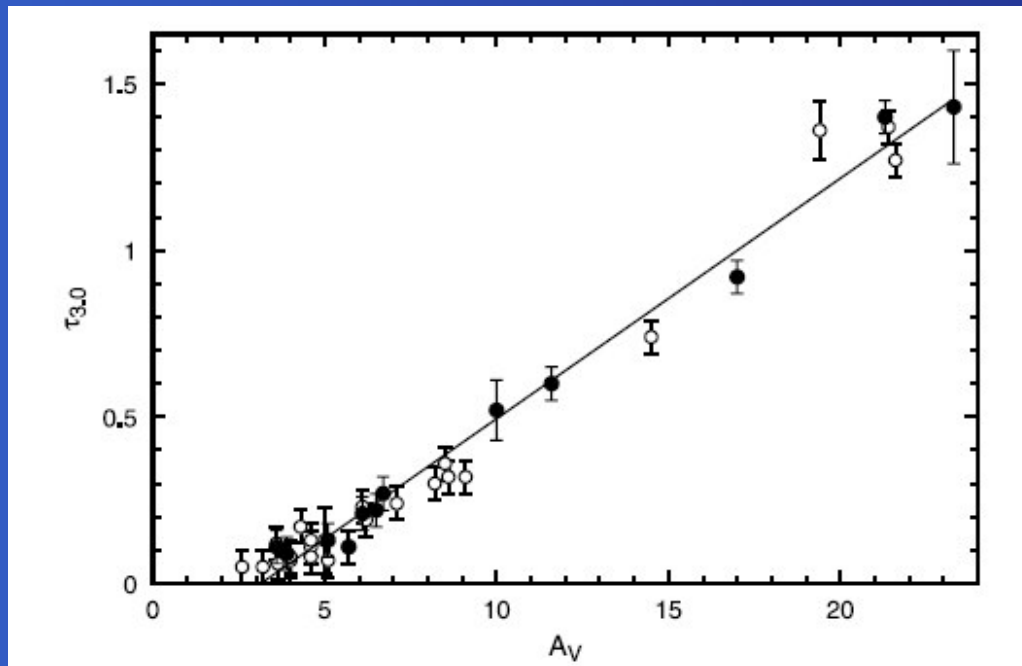


Whittet, ApJ (2001) 547, 872

Threshold value of  $A_V = 3$

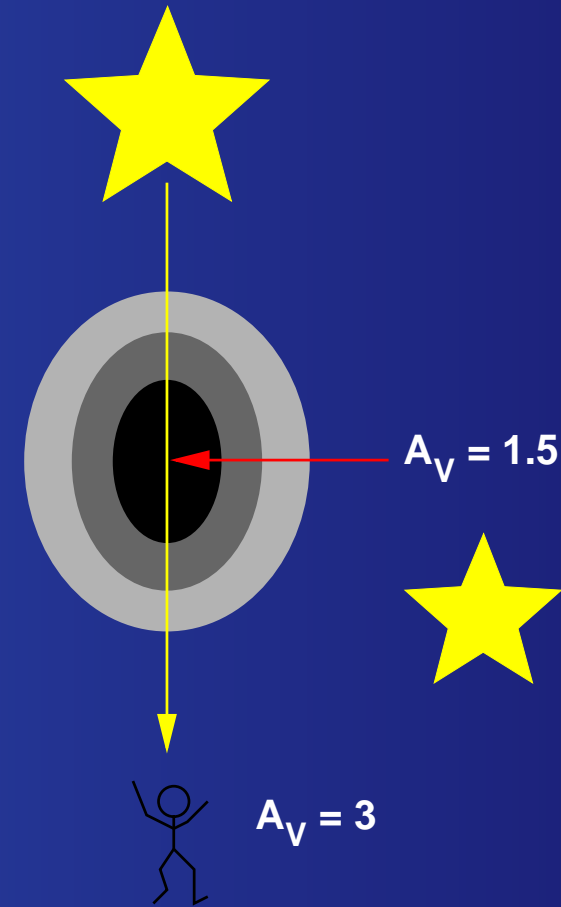
# Interstellar ice vs extinction

$$I_V(x) = I_V(0)10^{-A_V/2.5}$$



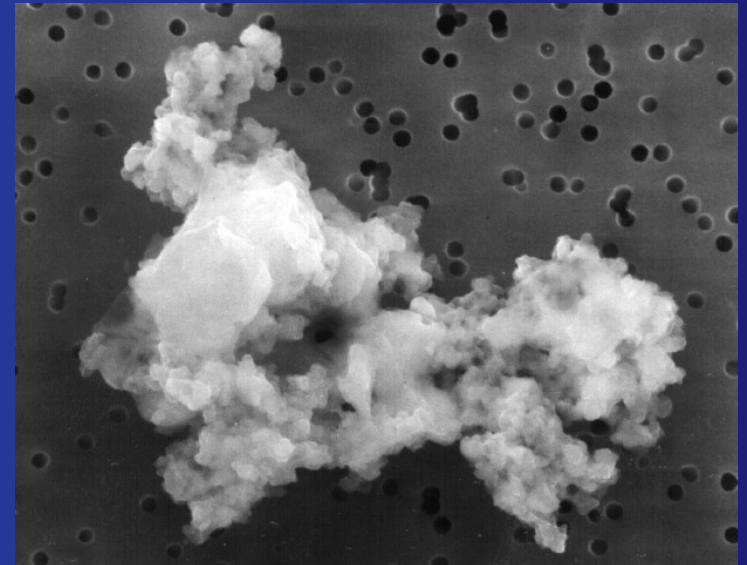
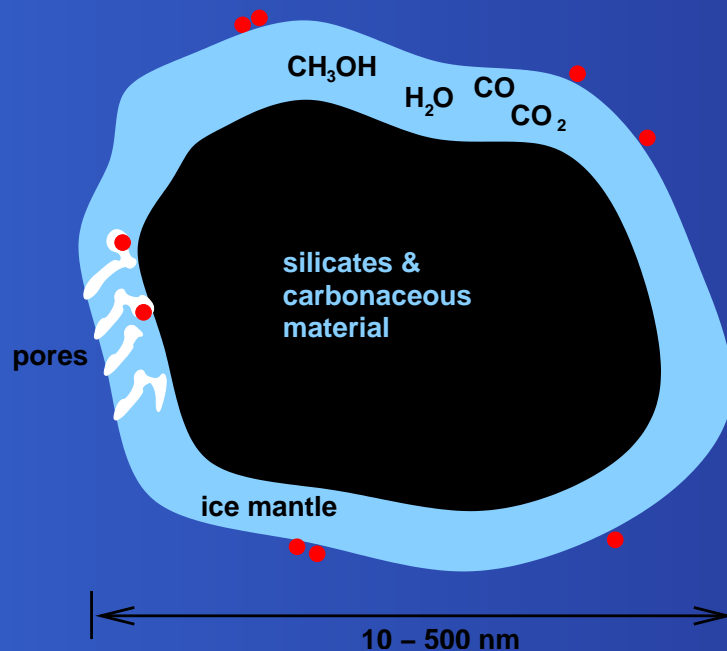
Whittet, ApJ (2001) 547, 872

Threshold value of  $A_V = 3$

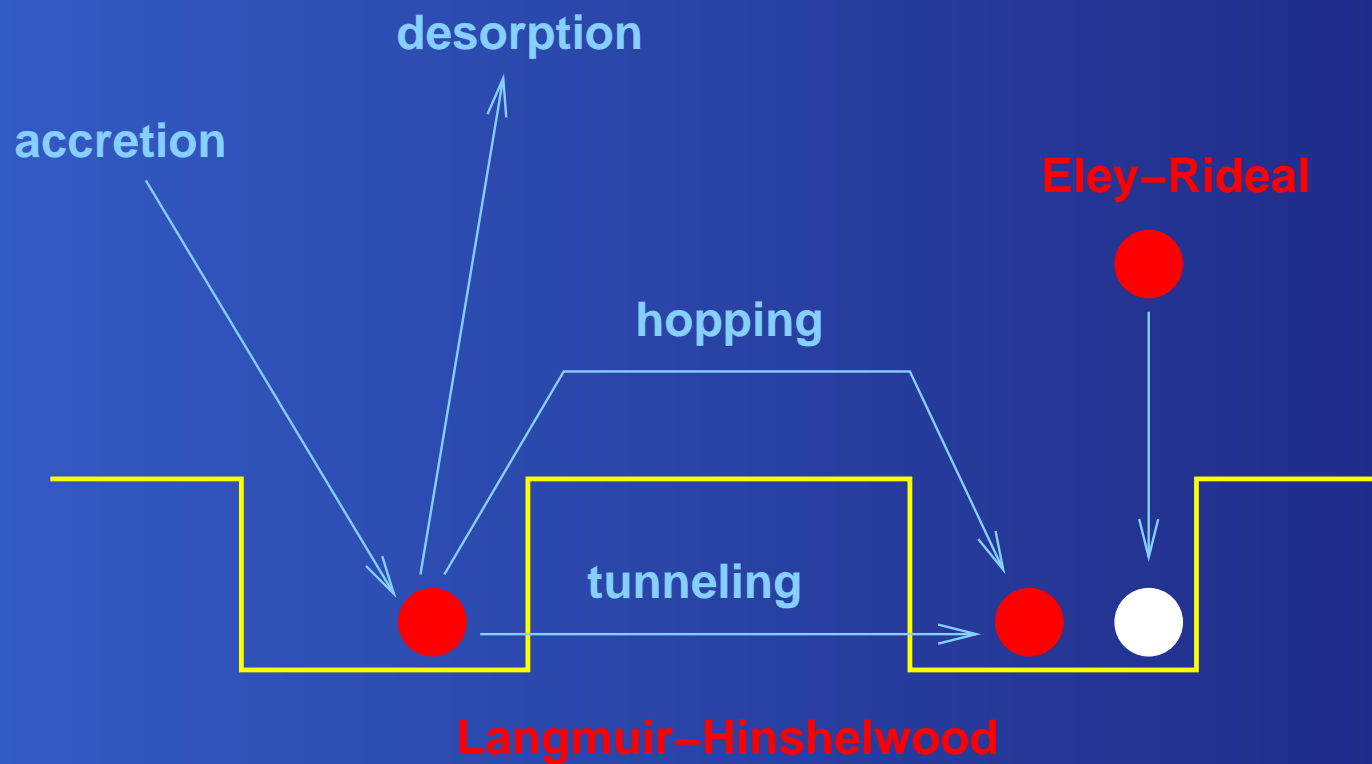


# Interstellar grains

- Ice present in dense and dark areas
  - Formed by surface reactions
  - Ice has porous character
- "Fluffy" shape



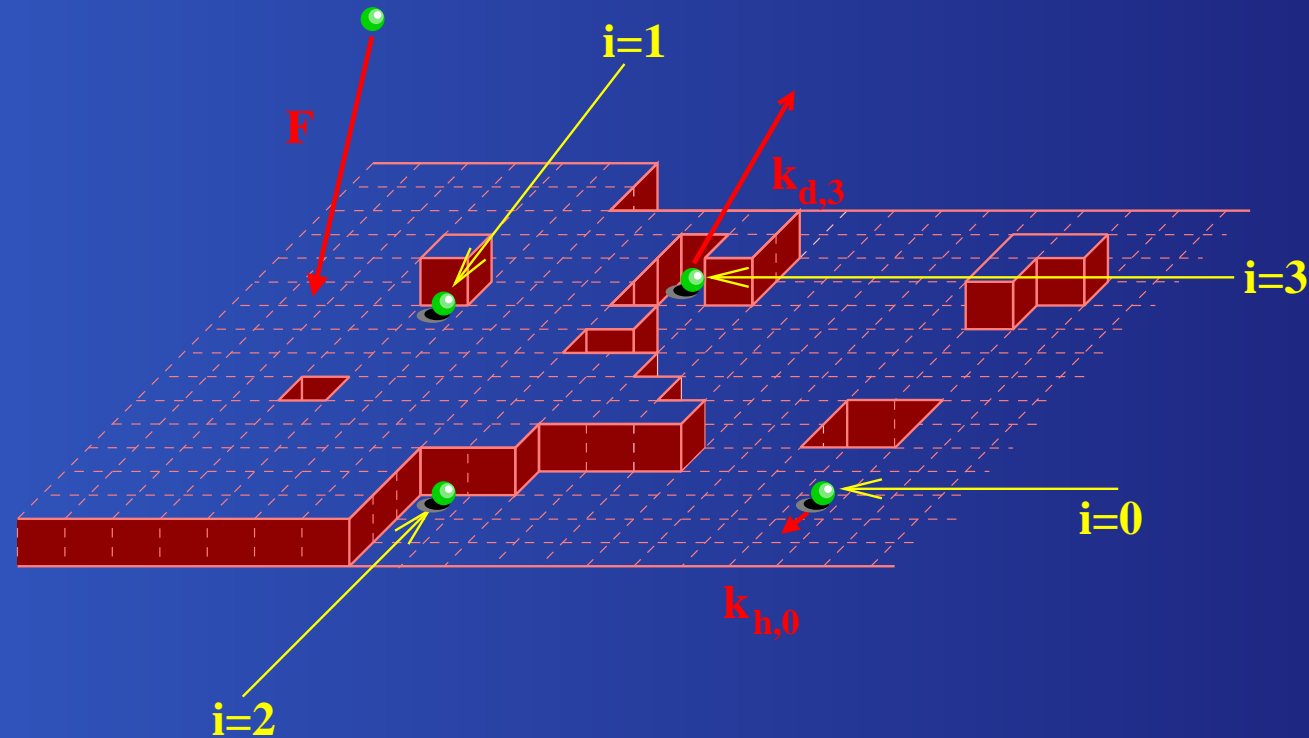
# Surface chemistry



"Physisorption"

"Flat"

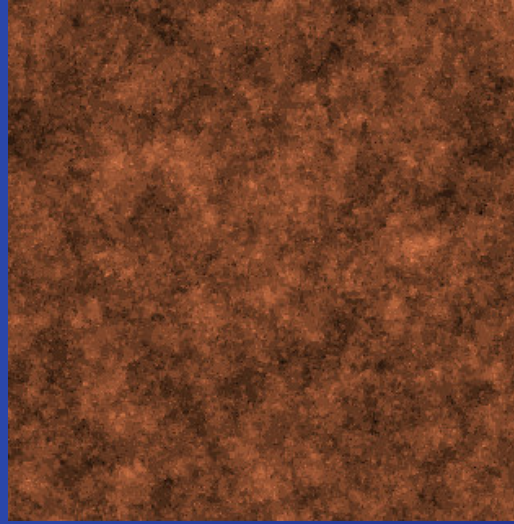
# Monte Carlo simulations



- Surface structure can be included
- Different rates depending on the number of "surface" neighbors
- Individual atoms can be followed



# Surface



$$k_{hop}^A = \nu \exp \left( - \frac{0.78 E^A + \alpha i_c E_c^A + \alpha i_{H_2O} E_{H_2O}^A}{kT} \right)$$

$$k_{eva}^A = \nu \exp \left( - \frac{E^A + \alpha i_c E_c^A + \alpha i_{H_2O} E_{H_2O}^A}{kT} \right)$$

# Surface reactions

Reaction			$\mu^1$	$E_a$ (K)
H	+	H $\rightarrow$ H <sub>2</sub>	0.991	0
H	+	O $\rightarrow$ OH	0.991	0
H	+	OH $\rightarrow$ H <sub>2</sub> O	0.991	0
O	+	O $\rightarrow$ O <sub>2</sub>	0.991	0
H	+	O <sub>2</sub> $\rightarrow$ O <sub>2</sub> H	0.991	1200
H	+	O <sub>2</sub> H $\rightarrow$ H <sub>2</sub> O <sub>2</sub>	0.991	0
H	+	O <sub>3</sub> $\rightarrow$ O <sub>2</sub> + OH	1	450
H	+	H <sub>2</sub> O <sub>2</sub> $\rightarrow$ H <sub>2</sub> O + OH	1	1400
H <sub>2</sub>	+	OH $\rightarrow$ H <sub>2</sub> O + H	1	2600
O	+	O <sub>2</sub> $\rightarrow$ O <sub>3</sub>	0.991	0

<sup>1</sup> Kroes and Andersson, Proc. IAU symp 231, (2005) p. 427

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Reaction			$\mu^1$	$E_a$ (K)
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H	+	O $\rightarrow$ OH	0.991	0
H	+	OH $\rightarrow$ H <sub>2</sub> O	0.991	0
O	+	O $\rightarrow$ O <sub>2</sub>	0.991	0
H	+	O <sub>2</sub> $\rightarrow$ O <sub>2</sub> H	0.991	1200
H	+	O <sub>2</sub> H $\rightarrow$ H <sub>2</sub> O <sub>2</sub>	0.991	0
H	+	O <sub>3</sub> $\rightarrow$ O <sub>2</sub> + OH	1	450
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H	+	O <sub>3</sub> $\rightarrow$ O <sub>2</sub> + OH	1	450
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<sup>1</sup> Kroes and Andersson, Proc. IAU symp 231, (2005) p. 427

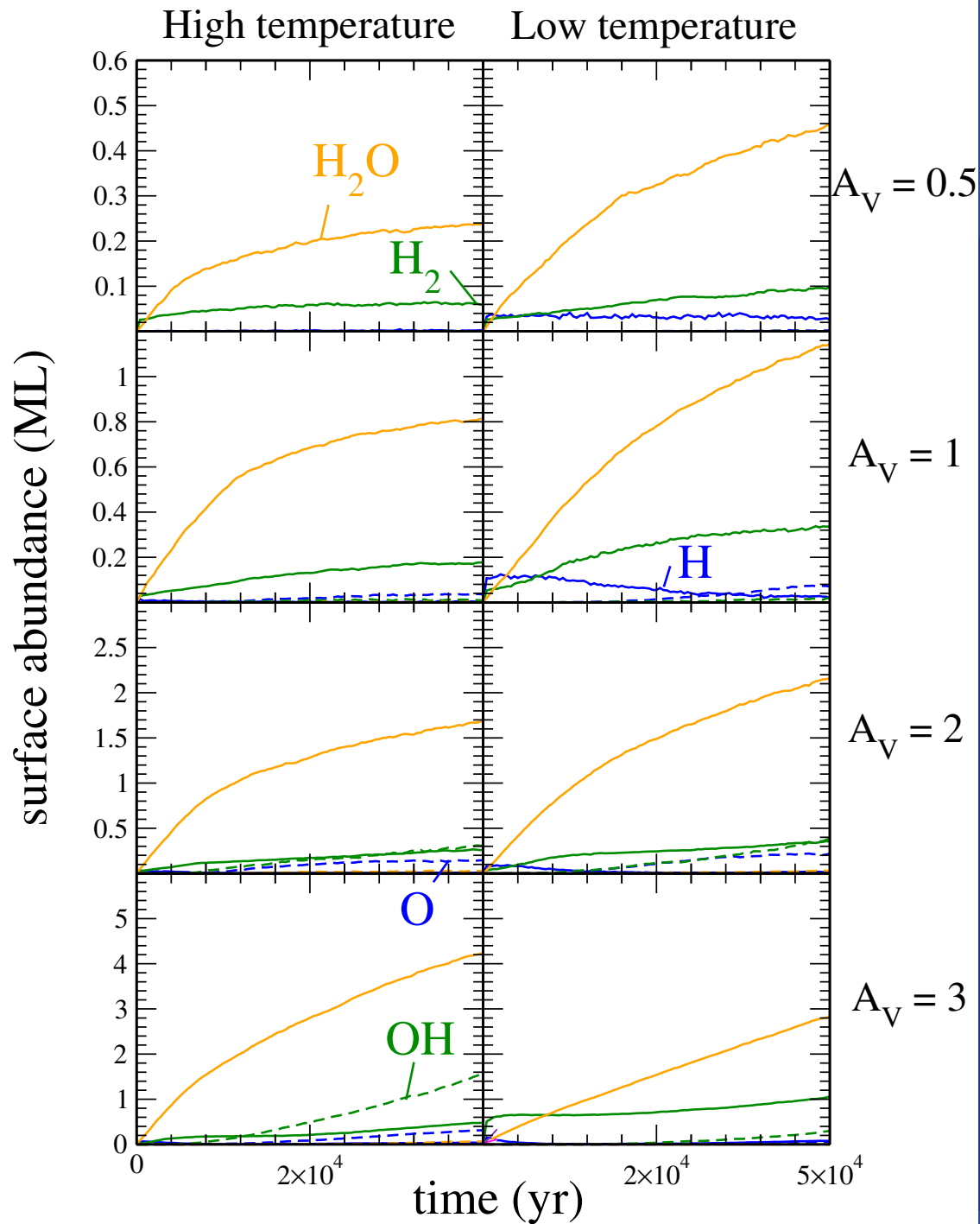
# Surface reactions

Reaction			$\mu^1$	$E_a$ (K)
H	+	H $\rightarrow$ H <sub>2</sub>	0.991	0
H	+	O $\rightarrow$ OH	0.991	0
H	+	OH $\rightarrow$ H <sub>2</sub> O	0.991	0
O	+	O $\rightarrow$ O <sub>2</sub>	0.991	0
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H	+	O <sub>3</sub> $\rightarrow$ O <sub>2</sub> + OH	1	450
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H <sub>2</sub>	+	OH $\rightarrow$ H <sub>2</sub> O + H	1	2600
O	+	O <sub>2</sub> $\rightarrow$ O <sub>3</sub>	0.991	0

<sup>1</sup> Kroes and Andersson, Proc. IAU symp 231, (2005) p. 427

# Dissociation reactions

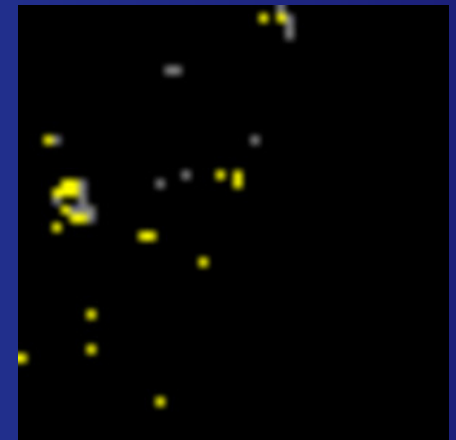
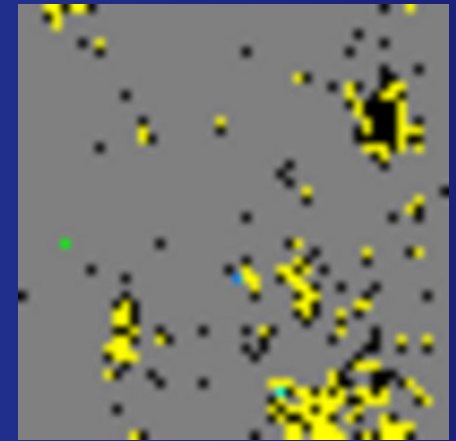
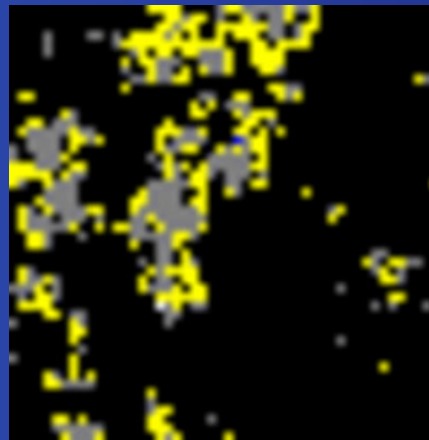
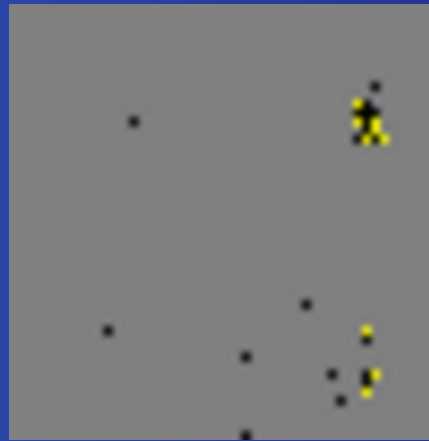
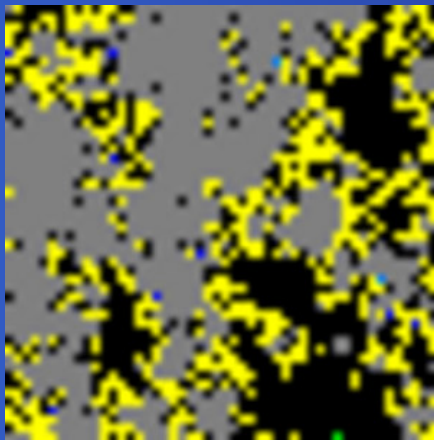
Reaction	$\alpha_{photo} \text{ (s}^{-1}\text{)}$	$\gamma_{photo}$	$\alpha_{CR}$
$\text{OH} \rightarrow \text{O} + \text{H}$	1.68(-10)	1.66	1.02(3)
$\text{H}_2\text{O} \rightarrow \text{H} + \text{OH}$	3.28(-10)	1.63	1.94(3)
$\text{O}_2 \rightarrow \text{O} + \text{O}$	3.30(-10)	1.4	1.50(3)
$\text{O}_2\text{H} \rightarrow \text{O} + \text{OH}$	0	0	1.50(3)
$\text{O}_2\text{H} \rightarrow \text{H} + \text{O}_2$	0	0	1.50(3)
$\text{H}_2\text{O}_2 \rightarrow \text{OH} + \text{OH}$	0	0	3.00(3)



Diffuse  
regions

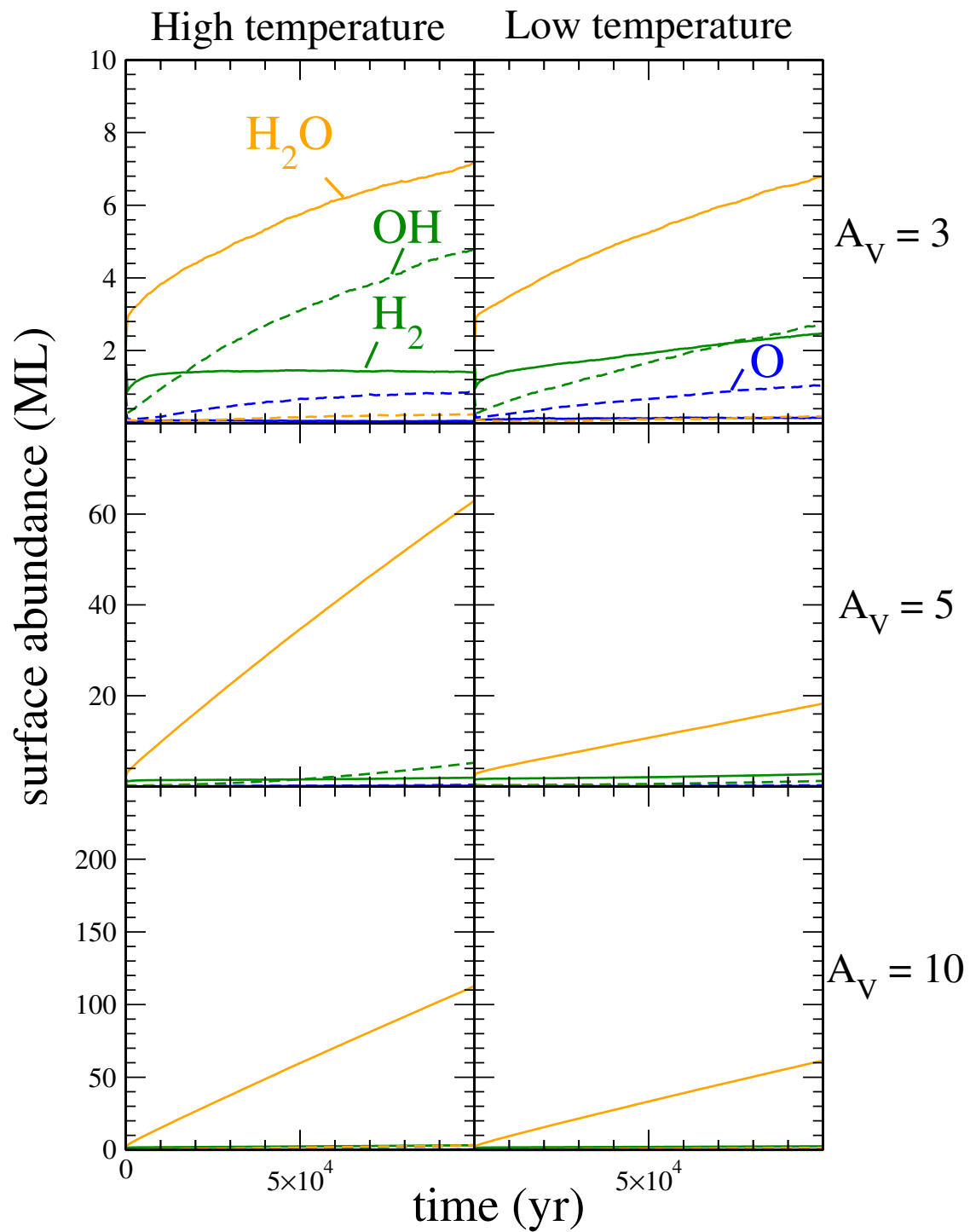
Threshold  
value  
 $A_V = 1.5$

# Ice surfaces

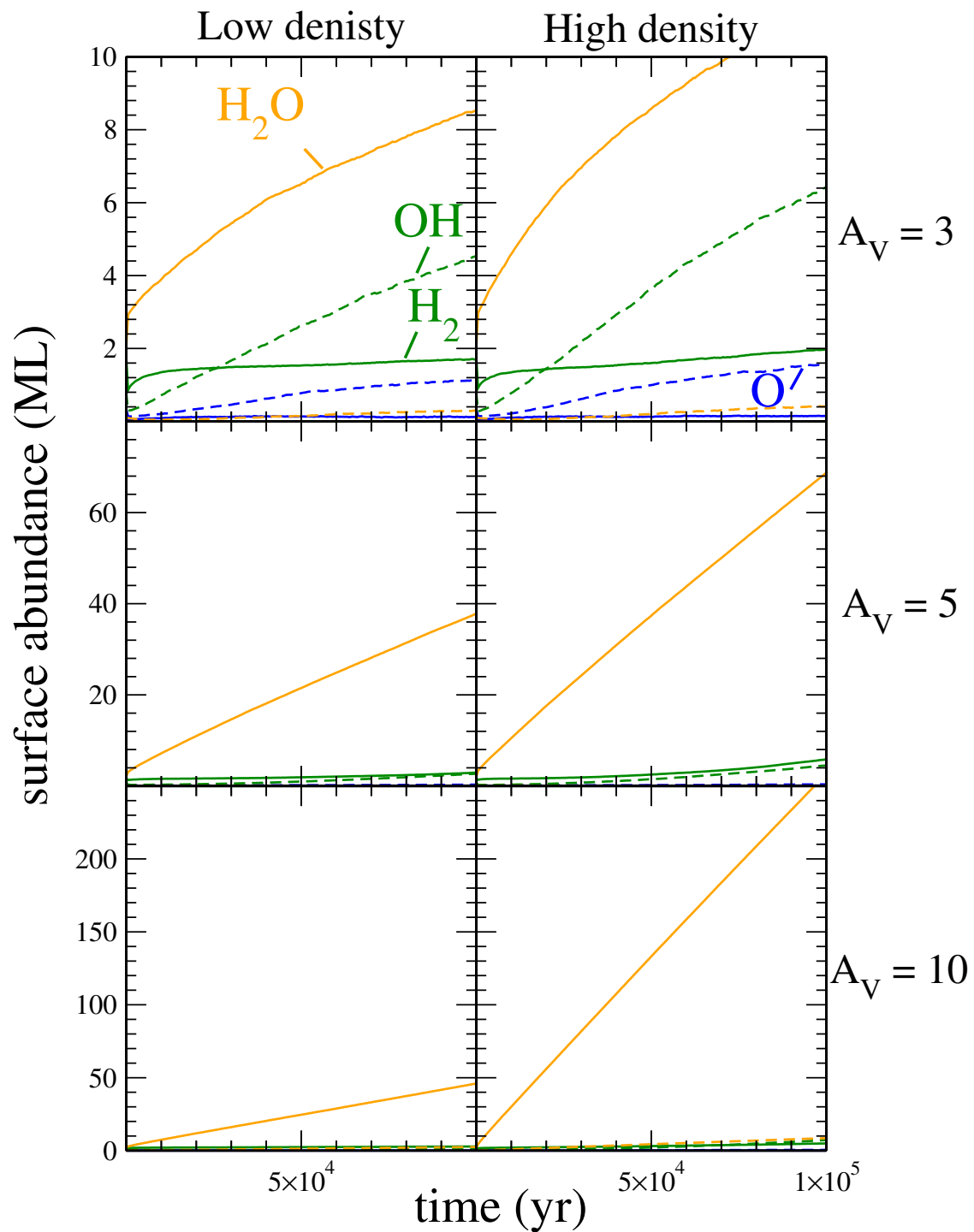


carbonaceous grain;  $\text{H}_2\text{O}$ ; ...;  $\text{H}$ ;  $\text{O}$ ;  $\text{OH}$ ;  $\text{H}_2$





Dense  
regions



Dense  
regions

# Conclusions

- Small surface coverage of ice in diffuse areas
- Ice function of temperature and density
- Ice mainly forms at surface steps
- Threshold value is in agreement
- H<sub>2</sub> blocks surface in dense areas

# Acknowledgments

- the Herbst group
- National Science Foundation for funding
- ... and you for your attention.

# Evaporation energies

Absorbate	Substrate	
	carbon	H <sub>2</sub> O
H	660	450
O	800	800
OH	1360	3500
H <sub>2</sub>	540	550
O <sub>2</sub>	1440	1000
H <sub>2</sub> O	2000	5640
O <sub>3</sub>	2240	1800
O <sub>2</sub> H	2100	1450
H <sub>2</sub> O <sub>2</sub>	2760	1900