## ROTATIONAL AND VIBRATIONAL ANALYSIS OF SOME LOW-LYING BENDING POLYADS IN THE $\tilde{A} {}^{1}A_{u}$ STATE OF ACETYLENE, C<sub>2</sub>H<sub>2</sub>

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A number of new low-lying vibrational levels of the  $\tilde{A}^{1}A_{u}$  state of acetylene have been identified, following infrared-ultraviolet double resonance experiments via the  $\nu_{3}$  ( $\Sigma_{u}^{+}$ ) and  $\nu_{3}+\nu_{4}$  ( $\Pi_{u}$ ) vibrational levels of the ground state, and high sensitivity one photon laserinduced fluorescence experiments with jet-cooled samples. These new levels involve the two lowest frequency vibrations, the torsion ( $\nu_{4}$ ) and the in-plane bend ( $\nu_{6}$ ), which are nearly degenerate and have been shown to be strongly coupled by *a*- and *b*-axis Coriolis interactions<sup>*a*</sup>. The most prominent bands in spectra recorded from the ground vibrational state or via  $\ell'' = 0$  vibrational intermediates go to  $K'_{a} = 1$  levels of the upper state (following the  $K'_{a} - \ell'' = \pm 1$  selection rule for the transition), however data from the  $K'_{a} \neq 0$ levels are affected by severe *a*-axis Coriolis coupling, which complicates vibrational assignment. Spectra recorded from  $\Pi_{u}$ -symmetry vibrational intermediate states access the  $K'_{a} = 0$  levels that reveal the purely vibrational interactions.

The combinations of  $\nu_3$  and  $\nu_6$  are highly anharmonic. Since theoretical calculations<sup>b</sup> indicate that the shape of the molecule at the *cis* – *trans* isomerization barrier will be a half-linear structure obtained by simultaneous excitation of  $\nu_3$  and  $\nu_6$ , this is not unexpected. The effective  $\nu_6$  interval in the highest assigned combination (3<sup>5</sup>6<sup>1</sup>) is found to have dropped to 60% of the fundamental frequency, indicating that it must lie close to the barrier.

<sup>&</sup>lt;sup>a</sup>A.L. Utz, J.D. Tobiason, E. Carrasquillo, M.L.J. Sanders and F.F. Crim, *J. Chem. Phys.* **95**, 2742 (1993). <sup>b</sup>J.F. Stanton, C.M. Huang and P.G. Szalay, *J. Chem. Phys.* **101**, 356 (1994).