## Millimeter and Submillimeter Wave Spectra of HNO<sub>2</sub> (*cis*), HNO<sub>2</sub> (*trans*), and HNO<sub>3</sub><sup>1</sup>

In gas phase equilibrium,  $HNO_2$  (nitrous acid) exists only as a component of a mixture which also contains NO,  $NO_2$ ,  $H_2O$ ,  $N_2O_3$ ,  $N_2O_4$ , and  $HNO_3$ . The resulting spectral overlap and reduced concentration have made studies of the rotational structure of  $HNO_2$  difficult, especially in the infrared. However, Allegrini *et al.* (1) have recently reported results from a CO laser intracavity Stark experiment, and high-resolution diode laser measurements are currently underway (2). The pure rotational structure of  $HNO_2$  has also been studied in the centimeter spectral region by  $Cox\ et\ al.$  (3) and Finnigan *et al.* (4).

In this paper we report measurements of 133 new millimeter and submillimeter wave transitions of HNO<sub>2</sub> and HNO<sub>3</sub>. The reported measurements and new analysis of HNO<sub>3</sub> are an extension of work previously reported by us (5). HNO<sub>2</sub> and HNO<sub>3</sub> are small asymmetric rotors with rotational constants of 10–100 GHz and large dipole moments. As such they exhibit an intense, dense spectrum throughout the millimeter and submillimeter wave spectral region. From the several thousand observable transitions, those reported have been selected so that they form the basis of analyses capable of producing accurate spectral maps for each species in this region of the spectrum.

TABLE I
Observed Transitions (MHz)

					W 754						
Transition	Strequency	CalcSbs.		requency	CalcObs.	Transition	observed Trequency	CalcObs.	Transition	Observed Frequency	0a1c0bs.
	HNG <sub>3</sub>		40, 37 - 40, 36 -4	57342.27	8 v. C-	163 13 = 162 1	311719.54	0.00	115 7 - 165 6	259086.74	0.00
			4J <sub>3-37</sub> - 39 <sub>3-36</sub> 5	44330.58	-35	171 16 - 170 1	7 270858.23	=0.000	11 <sub>4 3</sub> ~ 10 <sub>4 7</sub>	259206.31	0.20
18 <sub>7 11</sub> - 17 <sub>8 1</sub>	319514.60	-0.62	41 <sub>0 41</sub> - 40 <sub>1 40</sub> 9	19268.35	0.25	171 16 - 162 1	s 315434.00	0.39	11 <sub>4-7</sub> ~ 10 <sub>4-6</sub>	259210.56	0.05
20 <sub>5 15</sub> - 19 <sub>6 1</sub>	319382.89	0.01	41 <sub>2 39</sub> - 40 <sub>2 38</sub> 5	44309.07	-0.37	173 15 - 172 1	6 390007.14	-0.16	11 3 8 - 10 3 7	259607.46	0.02
21, 17 - 205 1	319346.13	3.05	42 <sub>1 41</sub> = 41 <sub>1 40</sub> 5	44283.39	-0.17	182 16 - 181 1	7 /06036.03	-0.01	11, 11 - 10, 10	254550.85	0.04
21 <sub>9 12</sub> - 20 <sub>10</sub>	11 182159.52	-0.08	43 <sub>0-43</sub> = 42 <sub>0-42</sub> 5	44252.97	0.72	19: 18 - 19: 1	321300.03	0.03	117 5 - 107 4	258954.10	0.01
223 19 ~ 214 1		0.03	4317 27 - 4317 26 -3		-0.03	192 17 - 191 1	8 217139.48	-0.01	11 <sub>8 4</sub> - 10 <sub>8 3</sub>	258912.20	0.00
232 21 - 223 2	319284.71	0.06	45 <sub>19-27</sub> - 45 <sub>19-26</sub> -3	29102.09	0.30	195 19 - 186 1	3 -319236.79	~0.08	11 <sub>9-3</sub> - 10 <sub>9-2</sub>	258874.87	0.01
241 23 - 232 2	319254.14	0.06	+8 <sub>22 27</sub> - 48 <sub>22 26</sub> -3	28142.95	0.11	20, 18 - 20, 1	3 230415.58	-0.04	11 <sub>10 2</sub> = 13 <sub>10 1</sub>	258839.26	-0.06
2411 13 - 2312	12 444783.24	-0.12	5023 28 - 5023 27 -3	40279.33	0.06	213 18 + 212 1	3 279196.57	3.07	12 <sub>2 13</sub> - 12 <sub>1 11</sub>	206817.18	-0.05
250 25 - 241 2	319221.53	0.09	50 <sub>24-27</sub> - 50 <sub>24-26</sub> -3	27420.52	0.13	21 <sub>5-16</sub> - 20 <sub>6-1</sub>	5 -268582.68	0.04	147 8 - 137 7	379591.10	0.00
25 <sub>5 20</sub> - 24 <sub>6 1</sub>	381908.39	0.02				222 20 - 221 2	1 263500.93	0.07	1410 5 - 1310 4	329405.00	0.03
2513 15 - 2411	14 444638.93	-6.10				223 19 - 222 2	0 275705.43	0.06	141 14 - 131 13	316907.59	0.01
26 <sub>9 17</sub> - 25 <sub>10</sub>	16 444556.47	-0.09	Cis HN	ic 2 at		233 20 = 232 2	1 273842.76	0.08	141 13 - 131 12	337198.86	0.02
278 19 - 269 1	8 444504.01	-0.02				23 <sub>5-18</sub> = 22 <sub>6-1</sub>	7 -217376.27	0.02	140 14 - 130 13	321236.27	0.02
272 26 - 273 2	s -319897.18	0.13	2.8 1.9	56957.24	0.05	231 22 - 230 2		-0.05	1411 - 1311 3	329353.07	0.03
270 27 - 260 2		0.08	1 9 3 9	56549.85	0.34	243-21 - 242-2		0.02	1412 3 - 1312 2	323301.07	-0.07
27, 27 - 27, 2	6 -332546.76	0.26	102 9 - 101 10 2	66316.39	0.03	- 25 <sub>2 23</sub> = 25 <sub>1 2</sub>	4 328250.73	0.06	160 16 - 150 15	365189.15	-0.01
28 <sub>7 21</sub> - 27 <sub>8 2</sub>	0 444467.19	0.00	1 10 0 0	74708.15	0.31	25 <sub>3-22</sub> - 25 <sub>2-2</sub>	3 275834.71	= G . U 5	161 15 - 151 14	384150.20	-9.01
29 <sub>6 23</sub> - 28 <sub>7 2</sub>	2 444438.62	0.01	112 10 - 111 11	76649.37	-0.01	253 22 - 244 2	1 203964.40	-0.02	16 <sub>2-14</sub> - 16 <sub>1-15</sub>	234696.36	-0.33
2914 15 - 2814	14 544866.66	-0.14	120 12 - 111 11 2	56148.69	0.01	- 16 <sub>3-23</sub> = 26 <sub>2-2</sub>	4 280028,24	-0.62	17 <sub>0 17</sub> - 16 <sub>0 15</sub>	387069.00	0.03
30 <sub>5-25</sub> - 29 <sub>6-2</sub>	444414.11	0.02	122 11 - 171 12 - 2	87945.39	0.02	27 4 24 - 272 2	5 286529.79	-0.05	182 16 - 181 17	211160.91	-0.12
30 <sub>13 17</sub> - 29 <sub>13</sub>	16 544685.06	G.00	123 9 = 122 10 3	38872.68	0.06	273 34 - 264 3	3 285170.03	0.02	190 19 - 180 18	430751.32	0.01
314 27 - 305 2		-0.32	5 13 1 12	83335.43	-9.04	276 21 - 267 2	-261791.16	0.06	19, 19 - 18, 18	428289.02	0.01
32 <sub>3 29</sub> - 31 <sub>4 2</sub>		-0.06	3 10 2 11	32960.86	0.03	286 23 - 387 2	-210345.14	-0.05	19, 15 - 18, 14	-277372.95	0.00
33 <sub>2 31</sub> - 32 <sub>3 3</sub>		-0.04	. 4 9 5 8	25162.96	0.04	303 27 - 302 2	8 320659.26	-0.01	202 18 - 201 19	223818.67	0.05
3310 23 - 3210		6.30	3 17 7 13	74551.54	-0.10	33, 29 - 32, 2	8 271920.51	0.00	30 <sub>4 16</sub> - 19 <sub>5 15</sub>	-252837.45	0.01
34 <sub>1 33</sub> - 33 <sub>2 3</sub>		-0.01	3 11 2 12	126345.89	0.00				211 20 - 210 21	333925.02	-0.03
347 28 - 347 2		-0.13	2 13 1 14	113374.36	0.03	Tra	ns HNO <sub>2</sub> ª		222 20 - 213 19	218966.68	0.01
<sup>35</sup> 0 35 - <sup>34</sup> 1 3		0.13	U 14 i 13	10011.10	-0.10		POR III		252 23 - 252 24	224897.25	-0.01
36 <sub>4 32</sub> - 35 <sub>5 3</sub>		-0.03	G 15 L 14	36167.60	-0.06	955 - 854	211944.29	0.04	26 <sub>5-21</sub> - 25 <sub>6-20</sub>	-271218.02	0.00
37 <sub>4 33</sub> - 36 <sub>5 3</sub>		-0.05	T 14 0 12	23686.00	0.03	9 <sub>6 4</sub> - 8 <sub>6 3</sub>	211899.86	0.63	285.23 - 276.22	-221734.75	0.00
<sup>38</sup> <sub>5</sub> <sub>33</sub> - <sup>37</sup> <sub>5</sub> <sub>3</sub>		0.06	2 14 1 15	327464.94	0.00	973 - 872		0.62	29 <sub>3-26</sub> - 28 <sub>4-25</sub>	217898.84	0.01
39 3 37 - 39 4 3		0.13	4 11 5 10	74913.52	-0.03	982 - 831	211838.93	0.04	313 28 - 312 29	323214.10	0.00
40 <sub>0 40</sub> - 39 <sub>1 3</sub>	9 506774.11	0.18	3 12 2 13	319193.49	0.01	11 <sub>6 6</sub> - 10 <sub>6 5</sub>		0.01	36 <sub>4-32</sub> - 36 <sub>5-31</sub>	235320.96	0.00
			1 15 2 14	81025.03	-0.01						
			162 15 - 161 16	342433.56	0.07						

a. Transitions have been corrected for the electric quadrapole interaction

<sup>&</sup>lt;sup>1</sup> This research is supported by NASA Grant NSG-7540.

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Rotational Constants (MHz)

	Cis	HNO <sub>2</sub>	Trans HNO	HNO2	HNO	
	value	Q	value	ρ	value	b
A	84 101.8381	0.011	92 892.0204	0.011	13 011.0485	0.0023
В	13 169.0755	0.0019	12 524.9742	0.0017	12 099.8803	0.0023
O	11 364.1644	0.0019	11 016.6671	0.0017	6 260.6488	0.0029
$\Delta_{J}(\times 10^{1})$	0.170502	0.000065	0.153081	0.000047	0.141390	0.000031
$\Delta_{ m JK}({ m \times}10^2)$	-0.86971	0.0076	5.17709	9 4 0 0 * 0	-2.018697	0.00033
$\Delta_{\mathrm{K}}(\times 10^{-1})$	0.195859	0.000028	0.299596	0.000053	.000738540	0.00000025
$\delta_{\rm J}({\rm x}10^2)$	0.280810	0.000023	0.206190	9 4 0 0 0 0 . 0	0.118256	0.000045
$\delta_{\mathrm{K}}(\mathrm{x}_{10}^{1})$	0.84811	0.00058	0.93120	0.0018	-0.205601	0,000048
H <sub>J</sub> (x10 <sup>7</sup> )	-0.3248	0.067	-0.2469	6+0.0	0.3145	0.012
H <sub>JK</sub> (×10 <sup>5</sup> )	-		-0.1735	0.024	-0.010756	0.00036
H <sub>KJ</sub> (×10 <sup>4</sup> )	-0.2027	0.011	-0.2297	0.0091	.0011465	0.00011
H <sub>K</sub> (×10 <sup>3</sup> )	0.1496	0.0086	0.2909	0.011	00003809	0.000008
h <sub>J</sub> (×10 <sup>8</sup> )	!	i i i	!	!!!	-0.9344	0.014
, h <sub>JK</sub> (×10 <sup>6</sup> )	-0.202	0.067	-0.439	0.19	-0.1369	0.0031
h <sub>K</sub> (x10 <sup>4</sup> )	0.4024	0.061	-0.2688	0.20	0.01089	0.00030
rms	6 # 0 • 0	6.1	8 40.0	8.	0.143	
Number of independent data points	ependent 73		99		175	

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Our basic millimeter and submillimeter wave experimental technique has been discussed previously (6). The cell used for most of the work was 1 m long, 2.5 cm diameter, and had Teflon windows at the ends. The inside of the cell was coated with clear Krylon to inhibit chemical decomposition. For *cis*-and *trans*-HNO<sub>2</sub>, we also utilized a 3-m cell.

HNO<sub>3</sub> is a stable molecule and a standard laboratory preparation was used to generate the sample. Various methods of preparing our HNO<sub>2</sub> sample were tried, but it was found that the highest yield was achieved by following the procedures of Varma and Curl (7). The final method adopted was to mix a few Torr each of NO<sub>2</sub>, NO, and H<sub>2</sub>O in the ratio of 2:1:1 in a mixing bulb. The sample was isolated and allowed to equilibrate for a few minutes. The mixing bulb was insulated and the temperature of the sample could be varied to maximize the concentration of HNO<sub>2</sub>. Measurements were made with a continuous flow of gas from the mixing bulb through the cell. Pressure in the cell was typically 0.05 Torr.

 $HNO_3$  and  $HNO_2$  are asymmetric rotors with electric dipole moments that give rise to both a- and b-type transitions. We have used Watson's reduced centrifugal distortion Hamiltonian (8) and the computational and statistical techniques that we have discussed previously (9, 10) for the analysis of the rotational spectrum of these molecules.

The high barriers to internal rotation divide the spectrum of HNO<sub>2</sub> into that of two distinct asymmetric rotors which are labeled *cis* and *trans*. Previous lower-frequency measurements (<40 GHz) (3, 4) were included in our analysis of both of these species. We measured 49 *cis* transitions and 40 *trans* transitions, extending the data set into the submillimeter region. The measurements were corrected for shifts due to an electric quadrupole interaction. Our data are presented in Table I along with the differences between these measured values and the frequencies calculated from the spectral constants shown in Table II. The analyses that lead to these spectral constants also include the earlier data discussed above. As can be seen from the rms deviations of the analyses, the fit to the earlier data is comparable to that shown in Table I.

For HNO<sub>3</sub> the data base consists in part the 111 transitions measured by Cazzoli and De Lucia in the millimeter range (5) and the 20 centimeter wave lines of Millen and Morton (11, 12). We have extended this data set by measuring 44 transitions in the submillimeter wave region of the spectrum. The results of our measurements are presented in Table II, and the new spectral constants are listed in Table II.

These new measurements and analyses provide accurate spectral maps throughout this spectral region. Because of their length they are on deposit at the Editorial Office of this journal. Anyone interested should first consult the second author and if unsuccessful should write to the Editor, Journal of Molecular Spectroscopy.

## REFERENCES

- 1. M. Allegrini, J. W. C. Johns, A. R. W. McKellar, and P. Pinson, J. Mol. Spectrosc. 79, 446–454 (1980).
- 2. A. Maki, private communication.
- 3. A. P. Cox, A. H. Brittain, and D. J. Finnigan, Trans. Faraday Soc. 67, 2179-2194 (1971).
- 4. D. J. FINNIGAN, A. P. Cox, A. H. BRITTAIN, AND J. G. SMITH, J. Chem. Soc. Faraday Trans. II 68, 548-565 (1972).
- 5. G. CAZZOLI AND F. C. DE LUCIA, J. Mol. Spectrosc. 76, 131-141 (1979),
- 6. P. Helminger, F. C. De Lucia, and W. Gordy, Phys. Rev. Lett. 25, 1397-1399 (1970).
- 7. R. VARMA AND R. F. CURL, J. Phys. Cham. 80, 402-409 (1976).
- 8. J. K. G. WATSON, J. Chem. Phys. 45, 1360-1361 (1966).
- 9. P. Helminger, R. L. Cook, and F. C. De Lucia, J. Mol. Spectrosc. 40, 125-136 (1971).
- 10. R. L. Cook, F. C. De Lucia, and P. Helminger, J. Mol. Spectrosc. 41, 123-136 (1972).
- 11. D. J. MILLEN AND J. R. MORTON, Chem. Ind. (N. Y.), 954 (1956).
- 12. D. J. MILLEN AND J. R. MORTON, J. Chem. Soc., 1523-1528 (1960).

WAYNE C. BOWMAN FRANK C. DE LUCIA

Department of Physics Duke University Durham, North Carolina 27706 PAUL HELMINGER

Department of Physics University of South Alabama Mobile, Alabama 36688

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