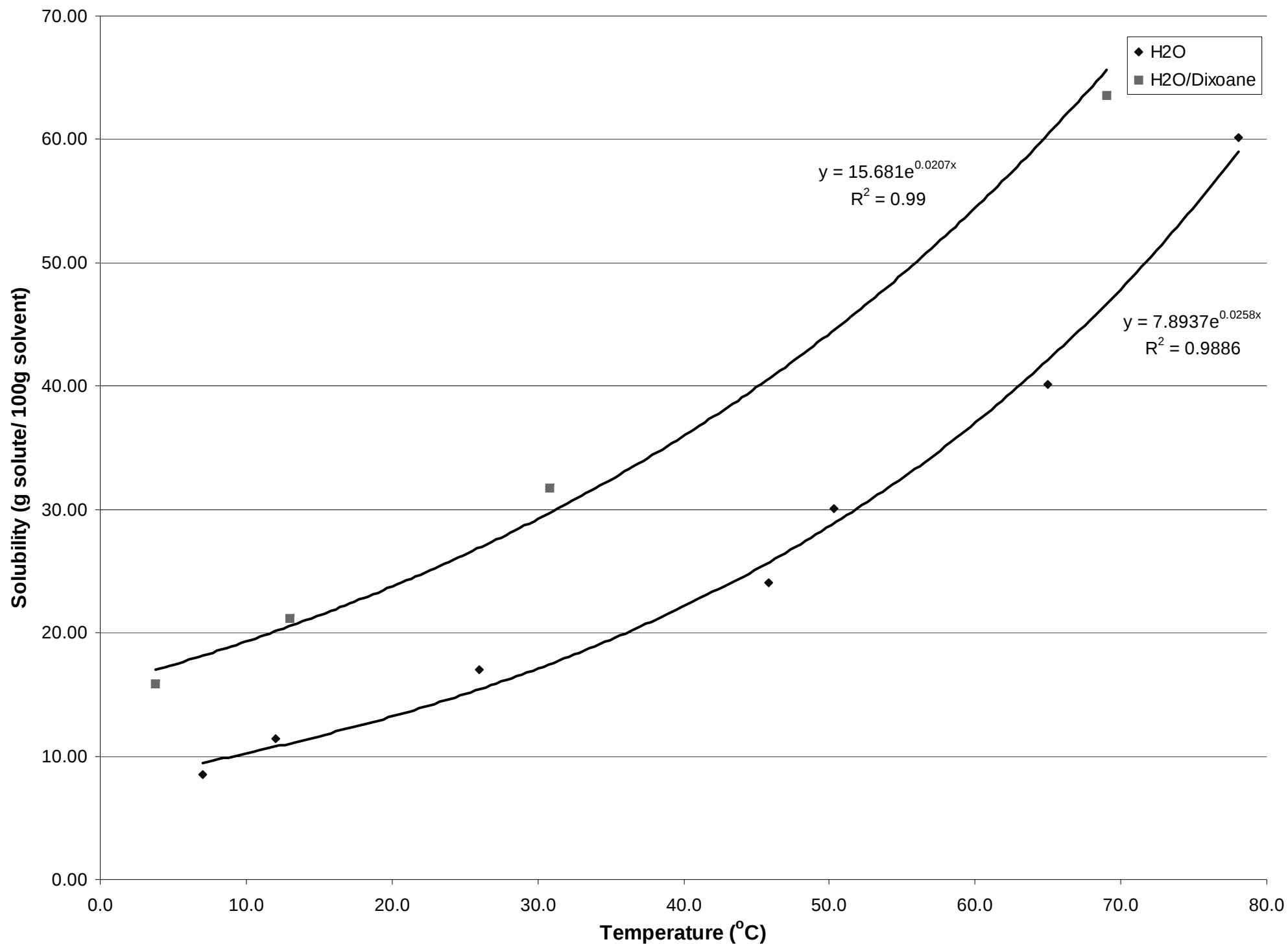


Exp 1: Variation of Solubility with Temperature and Solvent (SOL)						
Identity of solute: "Solute Name goes here"						
Day 1 Data: Water as solvent (data provided)			Mass of sample (g):	6.0003		
Addition	Total Volume (mL)	Temp at which crystals form (°C)	Concentration (g solute / 100 g solv.)	Mole Fraction		
1	10.00	78.1	60.15	0.07916		
2	15.00	65.0	40.10	0.05420		
3	20.00	50.3	30.08	0.04121		
4	25.00	45.8	24.06	0.03324		
A. Water as solvent: (data collected in lab)			Mass of sample (g):	1.7010		
Addition	Total Volume (mL)	Temp at which crystals form (°C)	Concentration (g solute / 100 g solv.)	Mole Fraction		
1	10.00	26.0	17.05	0.02379		
2	15.00	12.0	11.37	0.01599		
3	20.00	7.0	8.526	0.01204		

B. 70:30 water:dioxane as solvent:			Mass of sample (g):	6.5001		
Addition	Total Volume (mL)	Temp at which crystals form (°C)	Concentration (g solute / 100 g solv.)			
1	10.00	69.0	63.54			
2	20.00	30.8	31.77			
3	30.00	13.0	21.18			
4	40.00	3.8	15.88			

Solubility vs. Temperature for "solute name" in H₂O and H₂O/Dioxane Solvents



Temperature vs. Mole Fraction of "Solute Name" in H₂O

